

Solutions Worksheet 1 Molarity Complete The Table Answers

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Molarity Worksheet # 1
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Molarity Problems Worksheet
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Solutions Worksheet #1: Molarity Molar Mass of Mass of Molarity and Dilutions Worksheet - Google Docs
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78.9 g x 1 mole . Molarity = 303.76 g = 0.519 M 0.5000 L . Stoichiometry Worksheet # 3 . 1. Excess sodium hydroxide solution is added to 20.0 mL of 0.184 M ZnCl₂, calculate the mass of zinc hydroxide that will precipitate.

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NaCl in enough water to make 200.0 mL of solution. 4. Calculate the volume of 3.00 M copper(II) sulfate that must be diluted with water to prepare 500.0 mL of 2.15 M copper(II) sulfate. 5. Outline the correct laboratory procedure for diluting a 15.8 M stock solution of hydrochloric acid to a 1.0 L of a 3.0 M solution.
.00625 mol MgSO₄ (120g MgSO₄/1 mol MgSO₄) = .752 g MgSO₄, 250L (.025 mol MgSO₄)

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Molarity Worksheet W 331 Everett Community College Student Support Services Program What is the molarity of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.

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5. 125 cm³ of solution contains 3.5 moles of solute. What is the molarity of the solution? ? g KNO₃ = 0.175 mol KNO₃ × 101.1 g KNO₃ 1 mol KNO₃ = 17.7 g KNO₃ M = 3.5 mol 0.125 L = 28 M 6. Which solution is more concentrated? Solution "A" contains 50.0 g of CaCO₃ in 500.0 mL of solution. Solution "B" contains 6.0 moles of H₂SO₄

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2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution. Molarity Worksheet W 331 - Everett Community College Solutions Worksheet 1 Molarity Complete The Table Answers solutions worksheet 1 molarity complete the table answers, as one of the most functional sellers here will totally be in the midst of the best options to review.

Molarity Worksheet # 1

Solutions Worksheet 1 Molarity Complete The Table Answers are only available in Kindle format - users of other ebook readers will need to convert the files - and you must be logged into your Amazon account to download them. Solutions Worksheet 1 Molarity Complete Calculate molarity if 25.0 mL of 1.75 M HCl diluted to 65.0 mL. Calculate

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Molarity Problems Worksheet $M=nV$ $n=$ # moles V must be in liters (change if necessary) 1. What is the molarity of a 0.30 liter solution containing 0.50 moles of NaCl? 2. Calculate the molarity of 0.289 moles of FeCl₃ dissolved in 120 ml of solution? 3. If a 0.075 liter solution c

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a)+What+is+the+molarity+(M)+of+the+solution?+++ Molar+mass+of+NaCl+=58.44g/mole+ Moles+of+NaCl

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Solutions Worksheet #1: Molarity Molar Mass of Mass of

Molarity Problems Worksheet Use M or mol/L as unit for molarity. Remember that 1 Liter = 1000 mL. Do not confuse M, L, and mL! Some problems ask for volume - by algebra, $V = n/M$. Some problems ask for number of moles - $n = V M$. 1. What

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is the molarity of a 0.30 liter solution containing 0.50 moles of NaCl? 2.

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Molar Concentration of Solutions 1. What is the molarity of a solution made by

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Solutions Worksheet #1: Molarity 1) Complete the table: Solute formula Molar Mass of solute (g/mole) Mass of Solute (g)
Moles of solute (mole) Molarity of solution (M) Volume of solution (L) a NaCl 1.0 1.0 b NaOH 117.0 4.0 c MgCl₂ 190.3 2.0 d
NaCl 292.5 0.5 e KBr 238.0 2.0 f NaCl 1.5 1.0 g NaCl 1.0 0.25

Name: Date: Molarity

Problem #2: What is the molarity of 245.0 g of H₂SO₄ dissolved in 1.000 L of solution? Solution: $MV = \text{grams} / \text{molar mass}$
(x) (1.000 L) = 245.0 g / 98.0768 g mol⁻¹ x = 2.49804235 M to four sig figs, 2.498 M If the volume had been specified as
1.00 L (as it often is in problems like this), the answer would have been 2.50 M, NOT 2.5 M.

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Name: Date: Molarity About Chemistry <http://chemistry.about.com> Complete the table for the following aqueous solutions

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Calculate molarity if 25.0 mL of 1.75 M HCl diluted to 65.0 mL. Calculate molarity by dissolving 25.0g NaOH in 325 mL of
solution. Calculate grams of solute needed to prepare 225 mL of 0.400 M KBr solution. Calculate mL of 0.650M KNO₃
needed to contain 25.0g KNO₃. Which are water soluble? Zn(NO₃)₂ AlCl₃ AgBr FePO₄ CuAc₂

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